

WS 8.2-1 Writing Formulas for Ionic Compounds

1.	ZnBr ₂	11.	(NH ₄) ₂ Cr ₂ O ₇	21.	CuCl ₂	31.	Mn ₃ (PO ₄) ₄
2.	Na ₂ O	12.	AgCH ₃ COO	22.	Fe ₂ O ₃	32.	Fe(NO ₃) ₂
3.	LiOH	13.	Na ₂ Cr ₂ O ₄	23.	Mn(NO ₃) ₂	33.	CuCO ₃
4.	CaF ₂	14.	Li ₂ S	24.	PbBr ₄	34.	Zn(ClO ₃) ₂
5.	Ag ₂ S	15.	AlCl ₃	25.	Cr ₂ (CO ₃) ₃	35.	FeO
6.	(NH ₄) ₂ S	16.	Ca(NO ₃) ₂	26.	Sn(CrO ₄) ₂	36.	HgSO ₄
7.	MgC ₂ O ₄	17.	(NH ₄) ₂ O	27.	PbSO ₄	37.	PbS ₂
8.	BaSO ₄	18.	K ₂ S	28.	NH ₄ MnO ₄	38.	Fe ₂ (CO ₃) ₃
9.	KClO ₂	19.	Ag ₂ CO ₃	29.	Ag ₂ C ₂ O ₄	39.	K ₂ C ₂ O ₄
10.	Al(NO ₃) ₃	20.	Mg ₃ (PO ₄) ₂	30.	Fe(OH) ₃	40.	MnS

WS 8.3-1 Naming Ionic Compounds

1.	potassium chloride	21.	lithium oxide
2.	sodium sulphide	22.	sodium cyanide
3.	aluminum chloride	23.	silver chromate
4.	barium oxide	24.	calcium chlorate
5.	silver sulphide	25.	ammonium bicarbonate
6.	aluminum oxide	26.	zinc iodide
7.	lithium fluoride	27.	potassium permanganate
8.	zinc fluoride	28.	barium bromide
9.	magnesium bromide	29.	calcium phosphate
10.	calcium sulphide	30.	sodium dichromate
11.	potassium nitrate	31.	lithium nitrate
12.	magnesium sulphate	32.	magnesium sulphide
13.	zinc hydroxide	33.	sodium hypochlorite
14.	ammonium iodide	34.	potassium monohydrogen phosphate
15.	sodium carbonate	35.	calcium hydroxide
16.	magnesium hydrogen sulphate	36.	ammonium phosphate
17.	silver hydroxide	37.	aluminum dihydrogen phosphate
18.	zinc phosphate	38.	silver chloride
19.	ammonium sulphate	39.	potassium sulphite
20.	aluminum hydrogen sulphide	40.	sodium perchlorate

WS 8.3-2 Naming Ionic Compounds with Multivalent Metal Ions

1.	iron(II) oxide	21.	calcium hydroxide
2.	tin(IV) sulphide	22.	chromium(III) chloride
3.	lead(II) sulphate	23.	chromium(II) carbonate
4.	chromium(III) sulphide	24.	silver sulphate
5.	copper(II) nitrate	25.	ammonium fluoride
6.	iron(III) sulphate	26.	iron(III) dichromate
7.	tin(II) fluoride	27.	lead(II) sulphide
8.	mercury(II) sulphate	28.	copper(II) permanganate
9.	copper(II) phosphate	29.	chromium(III) sulphate
10.	manganese(II) permanganate	30.	copper(II) fluoride
11.	iron(II) hydroxide	31.	chromium(III) hydrogen carbonate
12.	lead(IV) chromate	32.	iron(III) phosphate
13.	copper(I) chloride	33.	sodium sulphide
14.	manganese(IV) oxide	34.	lead(IV) chloride
15.	tin(II) oxalate	35.	mercury(II) nitrate
16.	iron(II) chlorate	36.	chromium(II) oxide
17.	mercury(I) bromide	37.	mercury(I) nitrate
18.	copper(II) hydrogen sulphide	38.	calcium oxalate
19.	manganese(IV) carbonate	39.	barium phosphate
20.	lead(IV) nitrite	40.	tin(IV) sulphate

BLM 8.4-1 Chemical Families Study Guide

Students' responses should be similar to the following sample. Students should complete a chart for each chemical family.

Chemical Family		Properties That Are Shared by All Members of the Family
Noble gases		<ul style="list-style-type: none"> tasteless, colourless, odourless gas non-reactive
Element Symbol	Atomic Number	Properties That Are Different from Those of Other Family Members
He	2	–less dense than air
Ne	10	–bright red emission spectrum –similar density to air
Ar	18	–less reactive than Kr, Xe, Ra
Kr	36	–brilliant green and orange spectral lines
Xe	54	–blue emission spectrum
Ra	86	–radioactive