

## Chemical names and formulas of ionic compounds

1. Write the name for each of the following compounds.

(a)  $\text{BeS}$  Beryllium sulfide

(b)  $\text{Hg}_3\text{N}_2$  Mercury nitride

(c)  $\text{Cu}(\text{NO}_3)_2$  copper nitrate

(d)  $\text{Ag}_2\text{O}$  Silver oxide

(e)  $\text{CoBr}_2$  cobalt bromide

(f)  $\text{Bi}_3(\text{PO}_4)_5$  Bismuth phosphate

(g)  $\text{CaF}_2$  calcium fluoride

(h)  $\text{Mn}_2\text{O}_3$  Manganese oxide

(i)  $\text{Cr}_2(\text{SO}_4)_3$  chromium sulfate

(j)  $\text{ZnCl}_2$  Zinc chloride

(k)  $\text{Ni}(\text{OH})_2$  Nickel hydroxide

(l)  $\text{K}_2\text{Cr}_2\text{O}_7$  Potassium dichromate

(m)  $\text{ScF}_3$  Scandium fluoride

(n)  $\text{NaI}$  sodium iodide

(o)  $\text{Pb}(\text{CO}_3)_2$  Lead carbonate

(p)  $\text{RbClO}_2$  Rubidium chlorite

(q)  $\text{K}_3\text{P}$  Potassium phosphide

(r)  $\text{Mg}(\text{CN})_2$  Magnesium cyanide

(s)  $\text{SnS}$  tin sulfide

(t)  $\text{NaHCO}_3$  sodium hydrogen carbonate

2. Write the chemical formula for each of the following compounds.

(a) aluminum bromide  $\text{AlBr}_3$

(b) platinum(II) sulphide  $\text{Pt}(\text{SO}_3)_2$

(c) strontium sulfite  $\text{Sr}(\text{SO}_3)$

(d) scandium oxide  $\text{Sc}_2\text{O}_3$

(e) titanium(IV) nitride  $\text{Ti}(\text{NO}_2)_4$

(f) ammonium sulphate  $(\text{NH}_4)_2(\text{SO}_4)$

(g) lithium selenide  $\text{Li}_2\text{Se}$

(h) lead(II) hydrogen sulphate  $\text{Pb}(\text{HSO}_4)_2$

(i) sodium acetate  $\text{NaCH}_3\text{COO}$

(j) cesium chloride  $\text{CsCl}$

(k) cadmium(II) hydroxide  $\text{Cd}(\text{OH})_2$

(l) zinc phosphate  $\text{Zn}_3(\text{PO}_4)_2$

(m) barium chloride  $\text{BaCl}_2$

(n) tin(II) permanganate  $\text{Sn}(\text{MnO}_4)_2$

(o) lithium hypochlorite  $\text{LiClO}$

(p) gold(III) sulphate  $\text{Au}_2(\text{SO}_4)_3$

(q) sodium nitrate  $\text{NaNO}_3$

(r) chromium(III) chloride  $\text{CrCl}_3$

(s) potassium carbonate  $\text{K}_2\text{CO}_3$

(t) iron(III) bisulphate  $\text{Fe}(\text{HSO}_4)_3$