

Name _____

Date _____

Use with textbook pages 76–80.

Words to know about compounds

Vocabulary	
atom	ionic lattice
chemical bonds	lose
compound	molecule
covalent compounds	negatively
electrons	neutrons
element	polyatomic ion
gain	positively
ion	protons
ionic compounds	

Use the terms in the vocabulary box to fill in the blanks. Each term may be used more than once. You will not need to use every term.

- A pure substance that is made up of one type of atom is called a(n) _____.
- A pure substance that is made up of two or more types of atoms that are joined together due to a chemical change is called a(n) _____.
- Atoms in a molecule and ions in an ionic lattice are held together by _____.
- Chemical bonds are formed when atoms gain or lose _____ or when they share _____.
- When an atom loses electrons it becomes _____ charged.
When an atom gains electrons it becomes _____ charged.
- Metals and non-metals may form _____.
- The atoms in non-metals tend to _____ electrons.
- A(n) _____ is a repeating pattern of positive and negative ions.
- _____ form when non-metal atoms bond together by sharing their electrons.
- A neutral particle that is made up of atoms that are joined together by covalent bonds is called a(n) _____.
- A(n) _____ is an ion that is made up of two or more atoms that are held together with covalent bonds.

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Use with textbook pages 84–92.

Writing names and formulas of ionic compounds

You can use the periodic table on page 202 to help you answer these questions.

1. Complete the following table. First, identify each ion and its charge. Then, give the formula and name for each ionic compound formed. The table has been partially completed to help guide you.

	Chloride <u>Cl⁻</u>	Fluoride _____	Oxygen _____
sodium <u>Na⁺</u>	NaCl sodium chloride		
magnesium _____			
calcium _____			

2. Write the names of the following compounds.

(a) KCl _____

(e) ZnS _____

(b) LiBr _____

(f) SrO _____

(c) BaF₂ _____(g) AlCl₃ _____(d) Ag₃P _____(h) Mg₂C _____

3. Write the chemical formulas for the following compounds.

(a) beryllium sulphide _____

(e) calcium sulphide _____

(b) silver oxide _____

(f) lithium nitride _____

(c) sodium bromide _____

(g) rubidium chloride _____

(d) zinc chloride _____

(h) germanium bromide _____